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(c) Screening effect : Higher is the screening effect on the outer most electrons causes less attraction from the nucleus and can be easily removed, which is leading to the lower value of Ionisation Potential

$$\text{Ionisation Potential} \propto \frac{1}{\text{Screening effect}}$$

(d) Penetration power of sub shells :

(i) Order of attraction of subshells towards nucleus (Penetration power) is - $s > p > d > f$

(ii) As subshell is more closer to nucleus so more energy will be required to remove e^- in comparison to p,d & f. Ex.

	Be	B
	1s ² , 2s ²	1s ² , 2s ² 2p ¹
Ionisation Potential	Be	> B

After losing one e^- , B attains electronic configuration of Be, so IInd ionisation potential of B is more than Be. IInd Ionisation Potential of B > Be

(e) Stability of half filled and fully filled orbitals :

(i) Half filled p,d,f or fully filled s², p⁶, d¹⁰, f¹⁴ are more stable than others so it requires more energy.

Ex.

N	O
1s ² , 2s ² 2p ³	1s ² , 2s ² 2p ⁴
II nd Ionisation Potential order is O < N	

Because of half filled p-orbitals in N, its ionisation energy (stability) is higher than O.

IInd ionisation potential order Na < Al < Mg

(ii) Because s-orbital in Mg is completely filled and its penetration power is also higher than p-orbital (Al). IInd ionisation potential order Mg < Al < Na

(2,8,1)	(2,8,2)	(2,8)
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Periodic variation of ionisation energy

(a) Variation in period among the representative elements: Ionisation energies generally increases along the period because in moving left to right in a period the effective nuclear charge per outermost electron increases while the corresponding principal quantum number remain same.

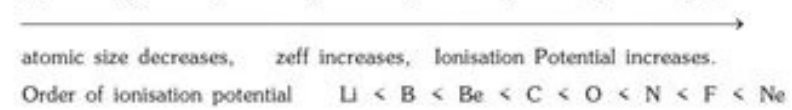
(b) Variation in a group among the representative elements : The ionisation energy generally decreases in moving from top to bottom because the size increases due to the increase of the principal quantum number. On the other hand the effective nuclear charge Z_{eff} for the outermost electron remains almost the same along the group.

Exception

(a) Ionisation Potential of Al < Ionisation Potential of Ga (While Ionisation Potential decreases down the group it is due to Transition contraction)

(b) Ionisation Potential of H< Ionisation Potential of Zr (While Ionisation Potential should decrease down the group, it is due to lanthanide contraction)

(c) In a period atomic size decreases and Z_{eff} increases so removal of electron becomes difficult and ionisation potential increases. But N, Be, P, Mg, show high ionisation energy than corresponding of next element.

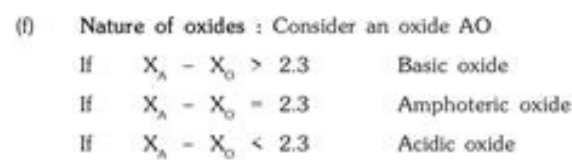
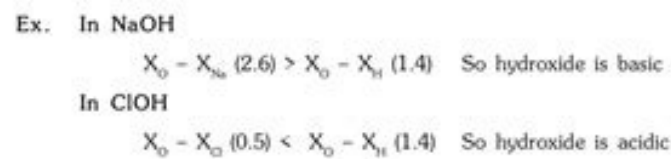
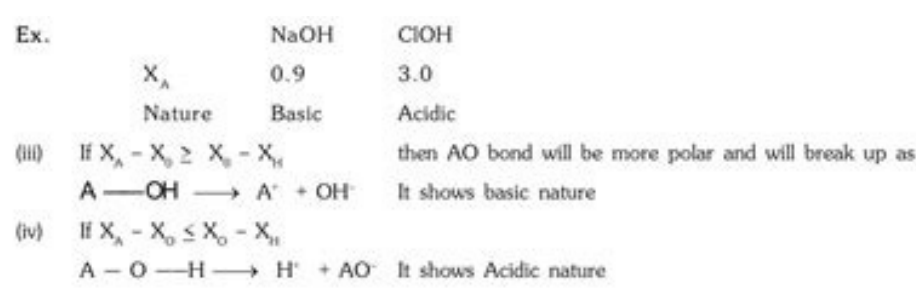


Periodic table of the elements

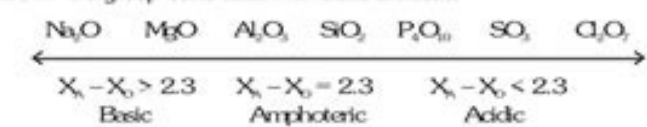
Periodic Table of the Elements

Periodic Table of the Elements

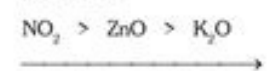
1. Period tables are called groups and they use 8 groups in Mendeleev's Periodic table.
2. The first group is called alkali metals and the second group is called alkaline earth metals.
3. The third group is called transition metals.
4. The fourth group is called transition metals.
5. The fifth group is called transition metals.
6. The sixth group is called transition metals.
7. The seventh group is called transition metals.
8. The eighth group is called transition metals.
9. The ninth group is called transition metals.
10. The tenth group is called transition metals.
11. The eleventh group is called transition metals.
12. The twelfth group is called transition metals.
13. The thirteenth group is called transition metals.
14. The fourteenth group is called transition metals.
15. The fifteenth group is called transition metals.
16. The sixteenth group is called transition metals.
17. The seventeenth group is called halogens.
18. The eighteenth group is called noble gases.
19. The lanthanide series is called lanthanides.
20. The actinide series is called actinides.



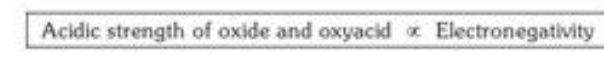
(g) Along a period acidic nature of oxide increases.
 (h) Down the group basic nature of oxide increases



ie. when in periodic table the distance between the element and oxygen increases, basic character increases.



Note: BeO, Al₂O₃, ZnO, SnO, PbO, SnO₂, PbO₂, Sb₂O₃ etc. are amphoteric oxides.
 CO, H₂O, NO, N₂O etc. are neutral oxides.



$SO_2 > SeO_2 > TeO_2$	$HNO_3 > H_2CO_3 > H_3BO_3$	$HClO_4 > HBrO_4 > HIO_4$
$H_2SO_4 > H_2SeO_4 > H_2TeO_4$	$H_3PO_4 > H_3AsO_4 > H_3SbO_4$	$N_2O_5 > P_2O_5 > As_2O_5$
$H_2SO_3 > H_2SeO_3 > H_2TeO_3$	$HOF > HOCl > HOBr > HOI$	$N_2O_3 > P_2O_3 > As_2O_3 > Sb_2O_3$

Acidic properties increases with increasing oxidation state of an element when central atom is same. Ex.
 (i) $Sb_2O_5 > Sb_2O_3$ (ii) $HClO_4 > HClO_3 > HClO_2 > HClO$

Mcq on periodic table pdf.

The concepts are elaborated in notes very precisely and to the point. In order to solve the given problems, aspirants are required to be well versed with all the formulas, named reactions, chemical equations, and periodic table - With the help of revision notes, they can easily grasp the knowledge of these. Chemistry is one of the most scoring subjects in JEE Examination and students should not leave any stone unturned for it to score high marks. These books are not enough. This book is mainly designed and written for the most of the questions falls in many entrance exams. You don't need many books for that and now when your concepts are well versed, you can go for MTG fingertips or any mcq sets/ model question papers. Data Storing Data storage refers to the use of recording media to retain data using computers or other devices. 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